

**Guanidinium Groups Act as General-Acid Catalysts in Phosphoryl Transfer Reactions: A Two-Proton Inventory on a Model System** [*J. Am. Chem. Soc.* **2004**, *126*, 9878–9879]. Anna M. Piątek, Mark Gray, and Eric V. Anslyn\*

Page 9878. We state, “(HPNP) is well accepted to undergo cleavage/transesterification by simple general-base catalysis.” However, it has been pointed out by Dr. Nicholas Williams from the University of Sheffield that early work on HPNP assigned the mechanism to specific-base catalysis (ref 13 of the paper). The mechanism, therefore, is not well accepted to involve general-base catalysis. This does not affect the conclusions from the two-proton inventory studies on compound **1**.

Page 9879. We reported a rate enhancement by the guanidinium in **1** to be 16 plus 26 to give a value of 42. As correctly pointed out by Dr. Nicholas Williams, these numbers should have been multiplied, giving a rate enhancement of around 420. These numbers assume that the isomeric hydroxypropyl substituents in **1** and HPNP have the same reactivity under the experimental conditions used.

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**Nature of Cl Bonding on the Au(111) Surface: Evidence of a Mainly Covalent Interaction** [*J. Am. Chem. Soc.* **2008**, *130*, 3720–3721]. Thomas A. Baker, Cynthia M. Friend, and Efthimios Kaxiras\*

Page 3720. An important reference was inadvertently omitted in the report. Migani and Illas<sup>1</sup> used density functional theory to systematically study the adsorption of halogens on transition-metal surfaces using bond distances, adsorption energies, work-function changes, vibrational frequencies, and Bader charge population analysis. Their observed periodic trends agree with our bonding analysis (done with charge density and density of states plots) that the bonding between Cl and the Au(111) surface is covalent in nature.

**Literature Cited**

(1) Migani, A.; Illas, F. *J. Phys. Chem. B* **2006**, *110*, 11894.

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